

Procedure

Using your computer, tablet or mobile device, navigate to the website:

<http://www.teachchemistry.org/bonding>. You should see the picture below on your screen.

Choose elements from the periodic table to bond.

H																		He
Li	Be										B	C	N	O	F			Ne
Na	Mg										Al	Si	P	S	Cl			Ar
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br		Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I		Xe
Cs	Ba	...	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At		Rn
Fr	Ra	...	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Uub	Uut	Fl	Uup	Lv	Uus		Uuo

Part 1: Ionic Bonding

1. Choose Sodium (Na).
 - a. What *type of element* is it?
 - b. How many valence electrons does it have?
2. Choose Fluorine (F).
 - a. What *type of element* is it?
 - b. How many valence electrons does it have?
3. Answer the question on the screen, "What type of bond is this combination likely to form?"
 - a. Circle: Ionic or Covalent?
 - b. Choose the appropriate number of atoms to make the bond. Record the number of each atom below:
4. Watch the final animation closely (it will play continuously).
 - a. Describe the change in the number of valence electrons in the atoms as the bond is successfully formed:
 - b. What does the positive (+) charge indicate (mention specific subatomic particles in your answer)?

- c. What does the negative (-) charge indicate (mention specific subatomic particles in your answer)?
- d. What is the final overall charge?
- e. Record the name and molecular formula for the compound below:



Reset the selected data using the reset symbol.

- 5. Choose Calcium (Ca).
 - a. What *type of element* is it?
 - b. How many valence electrons does it have?
- 6. Choose Chlorine (Cl).
 - a. What *type of element* is it?
 - b. How many valence electrons does it have?
- 7. Answer the question on the screen, "What type of bond is this combination likely to form?"
 - a. Circle: Ionic or Covalent?
 - b. Choose the appropriate number of atoms to make the bond. Record the number of each atom below:
- 8. Watch the final animation closely (it will play continuously).
 - a. Why were more than 2 total atoms needed to create this compound?
 - b. Explain what happened to the valence electrons in each atom.
 - c. What is the final overall charge?
 - d. Record the name and molecular formula for the compound below:

- e. Have you noticed a pattern between the charge of the ion and the number of valence electrons it has? Explain how you can predict the charge based on the number of valence electrons, or the location of the element on the periodic table.



Reset the selected data using the reset symbol.

9. Using a periodic table, complete the table below, then use the simulation to check each of your predictions:

Atom #1	Number of Valence Electrons	Prediction of charge	Atom #2	Number of Valence Electrons	Prediction of charge	Molecular Formula	Name of compound
Na			O				
K			F				
Mg			Cl				
Ca			N				
Al			S				

Part 2: Covalent Bonding

1. You will first investigate 5 *diatomic* molecules. Diatomic molecules are made up of 2 atoms.
 - a. Select 2 fluorine atoms. How many valence electrons are in each fluorine atom?
 - b. Is a fluorine atom a metal or a non-metal?
 - c. Did the combination of these atoms create a covalent or ionic bond?
 - d. How are the valence electrons organized to form a bond between these atoms?

e. How is this different from the ionic bonds formed in the previous part of the activity?

f. What shape does this molecule form?

2.

a. Select 2 oxygen atoms. How many valence electrons are in each oxygen atom?

b. Is an oxygen atom a metal or a non-metal?

c. Did the combination of these atoms create a covalent or ionic bond?

d. How are the valence electrons organized to form a bond between the atoms?

e. How is this bond different from the bond in the fluorine molecule in question 1?

f. What shape does this molecule form?

3. Make predictions in the following table. Once completed, check your answers using the simulation.

Lewis dot structure for single atom	Cl	S	N
Lewis dot structure for diatomic molecule (Cl ₂ , S ₂ , N ₂)			
Molecular formula			
Name of shape			

4. More than two atoms can also be combined to form a covalent molecule. These molecules may form different shapes and will also follow a particular naming system. Select the following combinations of atoms, and complete the rest of the table as you interact with the simulation:

1 st atom choice	2 nd atom choice	Predict Formula	Molecular Name	Shape
S	F			
N	Cl			
Cl	F			

Part 3: Critical thinking

1. What are the differences between ionic and covalent bonds? Be sure to refer to *valence electrons* in your response.
2. How is naming ionic and covalent compounds different? Use specific examples in your answer.
3. Based on your knowledge of ionic and covalent bonds, complete the missing portions of the following table:

Name	Formula	Ionic or Covalent?
Beryllium bromide		
	PF ₃	
Sulfur diiodide		
Strontium Phosphide		
	Cs ₃ N	
	H ₂ O	