

Name: Key
 Hour: _____ Date: 2018

Chemistry: Ionic Binary Compounds: Single-Charge Cations

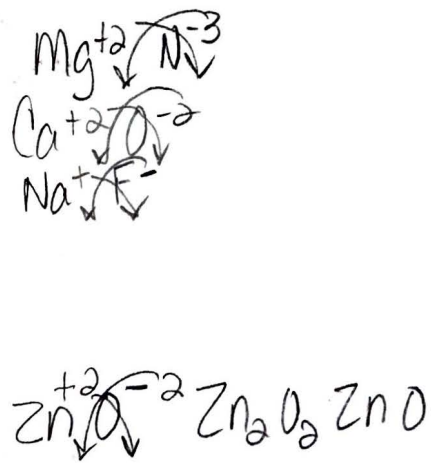
Write the name of each of the following compounds.

1. Na₂S
2. Al₂O₃
3. NaCl
4. RbI
5. ZnBr₂
6. AgCl
7. BN
8. BaF₂
9. Sr₃N₂
10. MgCl₂

1. Sodium sulfide
2. Aluminum oxide
3. Sodium chloride
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

Write the chemical formula for each of the given names.

11. magnesium nitride
12. calcium oxide
13. silver fluoride
14. beryllium chloride
15. potassium iodide
16. aluminum chloride
17. zinc oxide
18. barium bromide
19. lithium nitride
20. potassium sulfide



11. Mg₃N₂
12. CaO
13. NaF
14. _____
15. _____
16. _____
17. ZnO
18. _____
19. _____
20. _____

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Chemistry: Ionic Binary Compounds: Multiple-Charge Cations

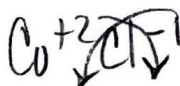
Write the name of each of the following compounds.

1. CuF
2. CuF₂
3. Cr₂O₃
4. PbI₂
5. PbCl₄
6. CrO₃
7. AuBr
8. NiO
9. V₂O₅
10. SnO₂

1. Copper (I) Fluoride
2. Copper (II) Fluoride
3. Chromium (III) Oxide
4. _____
5. _____
6. _____
7. _____
8. _____
9. _____
10. _____

Write the chemical formula for each of the given names.

11. manganese (VII) oxide
12. niobium (V) chloride
13. titanium (III) phosphide
14. palladium (IV) sulfide
15. platinum (II) fluoride
16. osmium (III) oxide
17. iridium (IV) nitride
18. cobalt (II) chloride
19. iron (III) sulfide
20. gold (III) iodide



11. Mn₂O₇
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. CoCl₂
19. _____
20. AuI₃