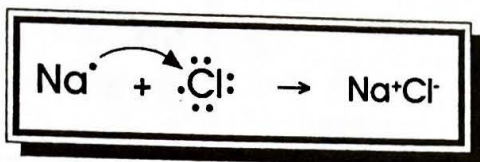


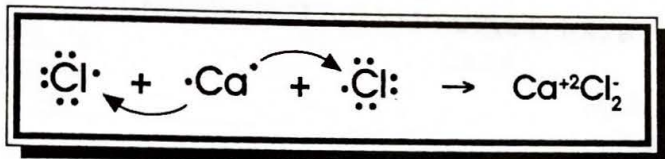
IONIC BONDING

Name Key

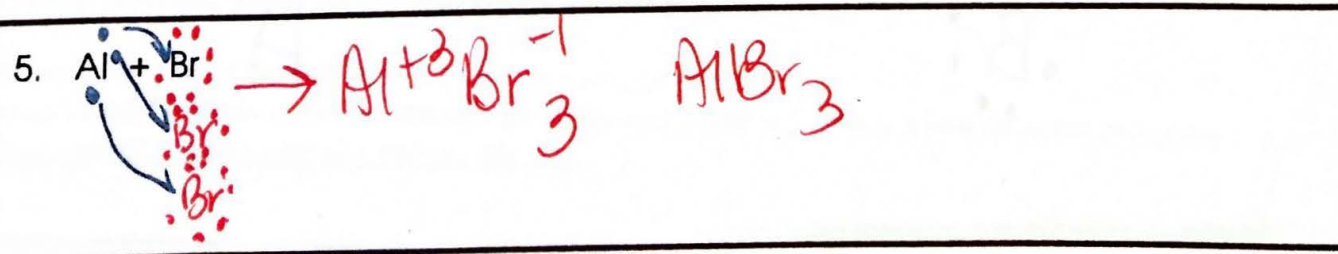
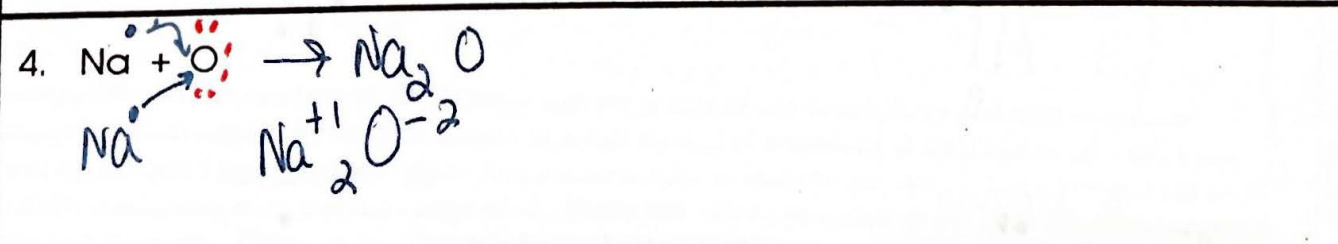
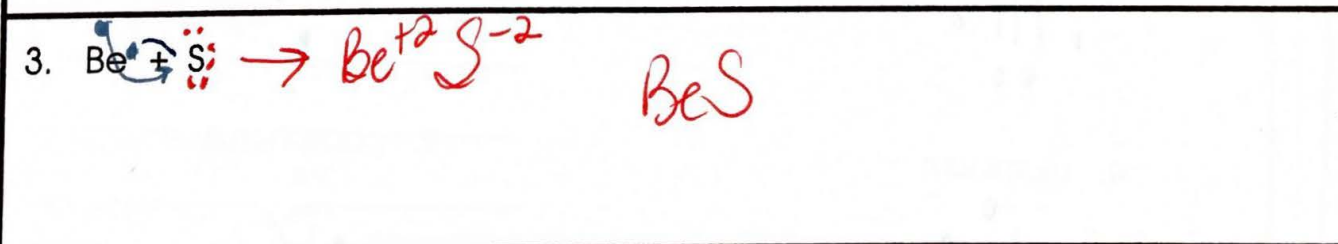
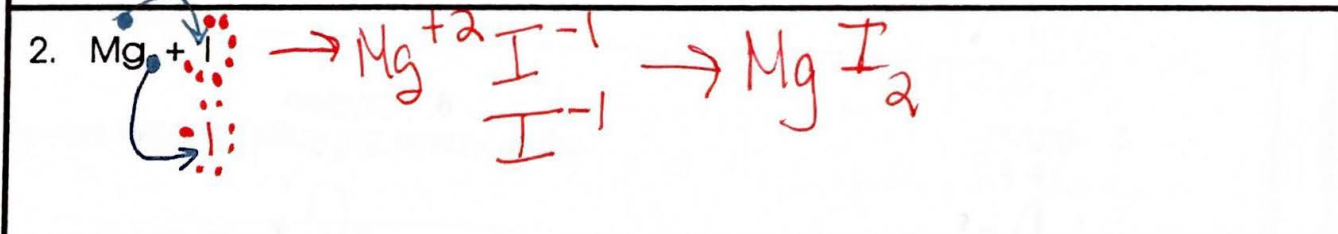
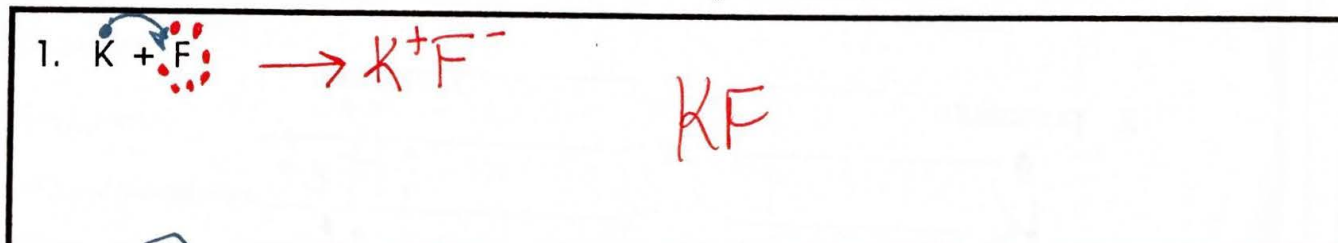
Ionic bonding occurs when a metal transfers one or more electrons to a nonmetal in an effort to attain a stable octet of electrons. For example, the transfer of an electron from sodium to chlorine can be depicted by a Lewis dot diagram.



Calcium would need two chlorine atoms to get rid of its two valence electrons.



Show the transfer of electrons in the following combinations.



LEWIS DOT DIAGRAMS

Name Kay

Lewis diagrams are a way to indicate the number of valence electrons around an atom.

Na^{\cdot} , $\cdot\ddot{\text{Cl}}\cdot$, $\cdot\ddot{\text{N}}\cdot$
are all examples of
this type of diagram.

Draw Lewis dot diagrams of the following atoms.

1. calcium



6. carbon



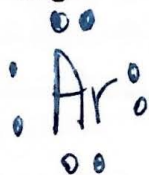
2. potassium



7. helium



3. argon



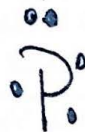
8. oxygen



4. aluminum



9. phosphorus



5. bromine



10. hydrogen

