

DETERMINING EMPIRICAL FORMULAS

Name _____

What is the empirical formula (lowest whole number ratio) of the compounds below?

1. 75% carbon, 25% hydrogen



2. 52.7% potassium, 47.3% chlorine



3. 22.1% aluminum, 25.4% phosphorus, 52.5% oxygen



4. 13% magnesium, 87% bromine



5. 32.4% sodium, 22.5% sulfur, 45.1% oxygen



6. 25.3% copper, 12.9% sulfur, 25.7% oxygen, 36.1% water



DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name _____

Key

Solve the problems below.

1. The empirical formula of a compound is NO_2 . Its molecular mass is 92 g/mol. What is its molecular formula?



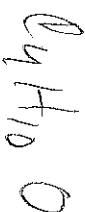
2. The empirical formula of a compound is CH_2 . Its molecular mass is 70 g/mol. What is its molecular formula?



3. A compound is found to be 40.0% carbon, 6.7% hydrogen and 53.5% oxygen. Its molecular mass is 60. g/mol. What is its molecular formula?



4. A compound is 64.9% carbon, 13.5% hydrogen and 21.6% oxygen. Its molecular mass is 74 g/mol. What is its molecular formula?



5. A compound is 54.5% carbon, 9.1% hydrogen and 36.4% oxygen. Its molecular mass is 88 g/mol. What is its molecular formula?

