in class 8/28/17

Name _____ Period _____ Date _____

Atomic Structure

PART A – SUBATOMIC PARTICLES

The table below contains information about several elements. In each case, enough information has been provided for you to fill in the blanks. **Assume all atoms are neutral.**

Isotope Name	Nuclear Symbol	Atomic Number	Mass Number	# of Protons	# of Electrons	# of Neutrons
1. calcium-40						
2.		12	24			
3.				1		2
4.	$^{197}_{79}Au$					
5.					26	30
6.			201	80		
7.		17				18

PART B – ANSWER THE FOLLOWING QUESTIONS:

8. What two particles, when added together, represent the Mass Number? _____ , _____

9. What is the Atomic Number equal to? _____

10. If an atom is Neutral, then what must be true? _____

11. What does the number represent in the Isotopes Name? (Ex. Calcium-40)

12. Can the Atomic Number be found on the periodic table? If so where? _____

13. Can the Atomic Mass be found on the periodic table? If so, where? _____

14. Can the Mass Number be found on the periodic table? If so, where? _____

15. Which particles are found in the nucleus? _____